

## ORGANIC CHEMISTRY

A Rh(II)-catalyzed multicomponent reaction by trapping an  $\alpha$ -amino enol intermediate in a traditional two-component reaction pathwayShunying Liu,<sup>1</sup> Wenfeng Yao,<sup>1</sup> Yuan Liu,<sup>1,2</sup> Qinghua Wei,<sup>1</sup> Jianghui Chen,<sup>1</sup> Xiang Wu,<sup>1</sup> Fei Xia,<sup>1,2\*</sup> Wenhao Hu<sup>1\*</sup>

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Multicomponent reactions (MCRs) represent an ideal organic synthesis tool for the rapid construction of complex molecules due to their step and atom economy. Compared to two-component reactions, the development of new MCRs has been greatly limited during the 170 years since the first MCR was reported. Theoretically, the trapping of an active intermediate generated from two components by a third component could change the traditional two-component reaction pathway, leading to the discovery of MCRs. We report an example of the trapping of  $\alpha$ -imino enols generated in situ from 1-sulfonyl-1,2,3-triazoles via  $\alpha$ -imino metal carbene species by vinylimine ions using C(2)-substituted indoles and paraformaldehyde as precursors in the presence of a rhodium(II) catalyst. The traditional enol-ketone transformation pathway was suspended by the trapping procedure and efficiently switched to an MCR pathway to produce  $\alpha$ -amino- $\beta$ -indole ketones in moderate to good yields. Unexpectedly, the resulting products and the theoretical density functional theory (DFT) calculation results indicated that the enolic carbon had a stronger nucleophilicity than the well-known traditional enamic carbon in the trapping process. The reaction mechanism was investigated using control experiments and detailed DFT calculations, and the synthetic application of the products was also illustrated. The developed strategy provides a mild and rapid access to  $\alpha$ -amino- $\beta$ -indole ketones and suggests a rationale for the discovery of MCRs by trapping an active intermediate with a third component in a traditional two-component reaction pathway.

## INTRODUCTION

Multicomponent reactions (MCRs) represent one of the most efficient approaches to realize atom-economical and green synthetic processes (1–8). Given the challenges in modulating the matching reactivity of all involved components, serendipity has always played an important role in the discovery of MCRs, even though several classical MCRs, such as the Strecker (9), Biginelli (10), and Ugi (11) reactions, are well established. Compared to two-component reactions, the development of new MCRs has been greatly limited during the 170 years since the first example of MCR was reported. Rational design strategies aimed at the discovery of novel MCRs have gained importance over the past decade (12–18). For instance, Nair *et al.* (14) demonstrated a conceptual framework involving the capture of a zwitterionic species by an electrically complementary component to develop a series of new MCRs for the construction of heterocyclic compounds. Other efficient rational design strategies, such as a single reactant replacement in a known MCR, the modulation of reaction sequences via versatile reactive intermediates, and the divergence of a reaction pathway by changing the reaction conditions, have been summarized in detail by Ruijter *et al.* (17). However, because of the continuous demand of novel MCRs, new strategies are urgently needed.

It is well known that numerous traditional two-component reactions proceed via an active intermediate through a sequence of two elementary steps, a reversible and an irreversible (rate-determining) step (19), in which the latter drives the equilibrium to the product side. When the irreversible step occurs after the generation of the active intermediate

(Fig. 1A, blue line), it is theoretically possible that the trapping of this active intermediate generated by two components with a third component might change the traditional reaction pathway and switch to a new energetically favored three-component pathway (Fig. 1A, red line).

Our group has continuously focused on the development of transition metal-catalyzed multicomponent transformations of diazo compounds, alcohols/amines/indoles, and electrophiles (20–28). In these transformations, diazo compounds and alcohols/amines/indoles generate in situ ylides or zwitterions that can be successfully trapped by electrophiles, such as imines and aldehydes. Whereas in the traditional reaction pathway, the active ylides or zwitterions readily undergo a proton transfer process (29–34) to form X–H (X = C, N, and O) insertion products. A theoretical DFT calculation investigation showed that it is the proton transfer step rather than the active ylide generation step that is the rate-determining step in the reaction, leading to the X–H insertion product formation (see fig. S1). Thus, we envisaged that the trapping of the active intermediate with a third component to suspend this irreversible two-component transformation step could provide a general strategy for discovering novel MCRs.

To validate this conceptual hypothesis, an MCR based on the trapping of another active intermediate with a third component beyond our present ylide or zwitterion systems is needed to be demonstrated: (i) By trapping an active two-component intermediate in an irreversible traditional pathway, the resulting three-component involved intermediate formed via an energetically favored transition state could be transformed into a multicomponent product; (ii) the multicomponent product formed by a stepwise cascade process from the reaction of the traditional two-component product with a third component might be excluded to ensure that the reaction is a “true” MCR according to the modern MCR definition (Fig. 1B) (4, 13).

After the seminal work of Gevorgyan *et al.* (35, 36), Fokin *et al.* (37–39), and others (40, 41), 1-sulfonyl-1,2,3-triazoles have attracted a great deal of attention as precursors of  $\alpha$ -imino Rh(II) carbene **A**, which is a practical intermediate in the synthesis of N-heterocyclic complexes. Recent additions

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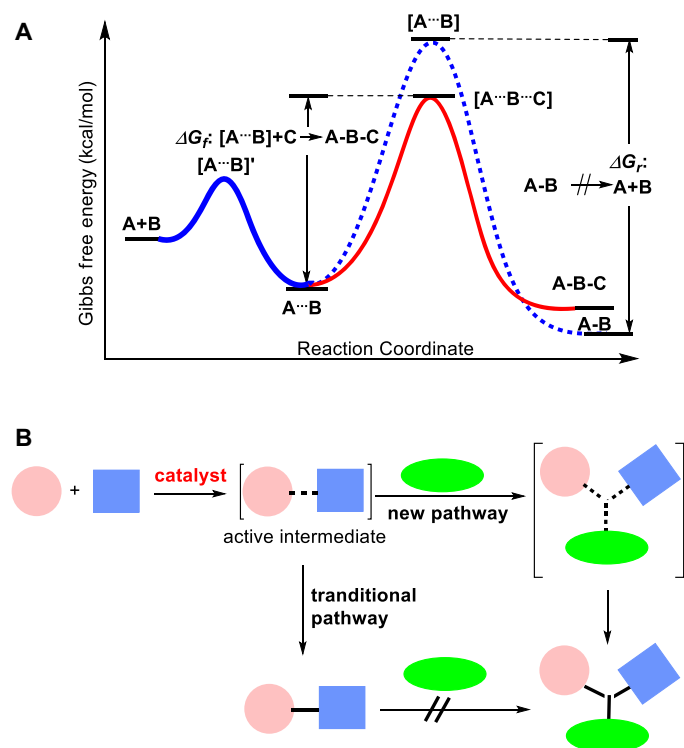
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to the rapidly growing list of applications of 1-sulfonyl-1,2,3-triazoles under Rh(II) catalysis include transannulations and cyclopropanations (42–49), C–H insertions (50), O–H insertions/rearrangements (51–55), ring expansions (56, 57), arylations with boronic acids (58), and cycloaddition reactions (59).  $\alpha$ -Amino enol **D** is a classically proposed key active intermediate via  $\alpha$ -imino Rh(II) carbene **A** from the reaction of 1-sulfonyl-1,2,3-triazole with water, as reported by Miura *et al.* (51). The authors proposed that carbene **A** underwent an O–H insertion with water to furnish intermediate **C**, which further yielded the key intermediate  $\alpha$ -amino enol **D** by imine-enamine tautomerization, and subsequently delivered an  $\alpha$ -amino ketone product by a subsequent keto-enol tautomerization via a proton transfer process (Scheme 1). Because the formation of **D** from  $\alpha$ -imino carbene **A** represents one of its most important transformations and provides a commonly accepted active intermediate (60, 61), the proposed  $\alpha$ -amino enol **D** was consequently selected as the substrate to be trapped to validate our hypothesis.

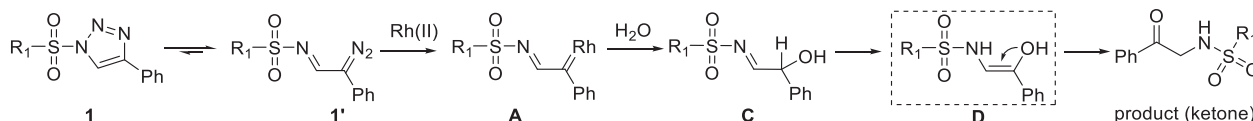
## RESULTS

### Preliminary exploration of the MCR

Initially, commonly used electrophilic trapping reagents such as aldehydes or imines were used to trap  $\alpha$ -amino enol **D**, as demonstrated



**Fig. 1. Designed strategy for the discovery of MCRs.** The rational design of a general strategy based on the trapping of the active intermediates in traditional two-component reactions for the discovery of MCRs. (A) Gibbs free energy profiles for two-component reactions and MCR. (B) Strategy for discovering MCR.



**Scheme 1. Proposed literature mechanism.** Proposed mechanism of the formation of ketones from triazoles via  $\alpha$ -amino enol **D** according to Miura *et al.* (51).

by Nair *et al.* (14). However, no formation of the desired products was observed when 4-(4-bromophenyl)-1-(methylsulfonyl)-1H-1,2,3-triazole **1** was reacted with water and aldehydes or imines using rhodium(II) acetate as catalyst. Only ketone **7** (~80% yield) and trace amounts of cycloaddition product **9** were obtained. These results suggested that a more active trapping component is needed to be taken into consideration to obtain a matching reactivity with  $\alpha$ -amino enol **D**. It has been widely reported that indoles react with aldehydes and ketones in the presence of Lewis or Brønsted acid to generate highly reactive vinyliimine ion **E** (Scheme 2A) (62–66), which undergoes an electrophilic reaction to yield functionalized indole derivatives. Owing to the broad application of indoles in drug discovery and our previous work on indole functionalization (21, 67–70), we thus turned our attention to the in situ-generated vinyliimine ion **E** as a trapping component to provide the multicomponent product **4** (Scheme 2A). For this designed reaction pathway, the most important competitive pathway is a two-component [3 + 2] cycloaddition that has been illustrated by Spangler *et al.* (47) for the efficient synthesis of pyrroloindoline architectures **5** via the combination of C(3)-substituted indole **2** and Rh(II)-associated 1-sulfonyl-1,2,3-triazoles (Scheme 2B). The authors suggested that the reaction was promoted via zwitterionic-type intermediate **F** from an  $\alpha$ -imino rhodium carbene species.

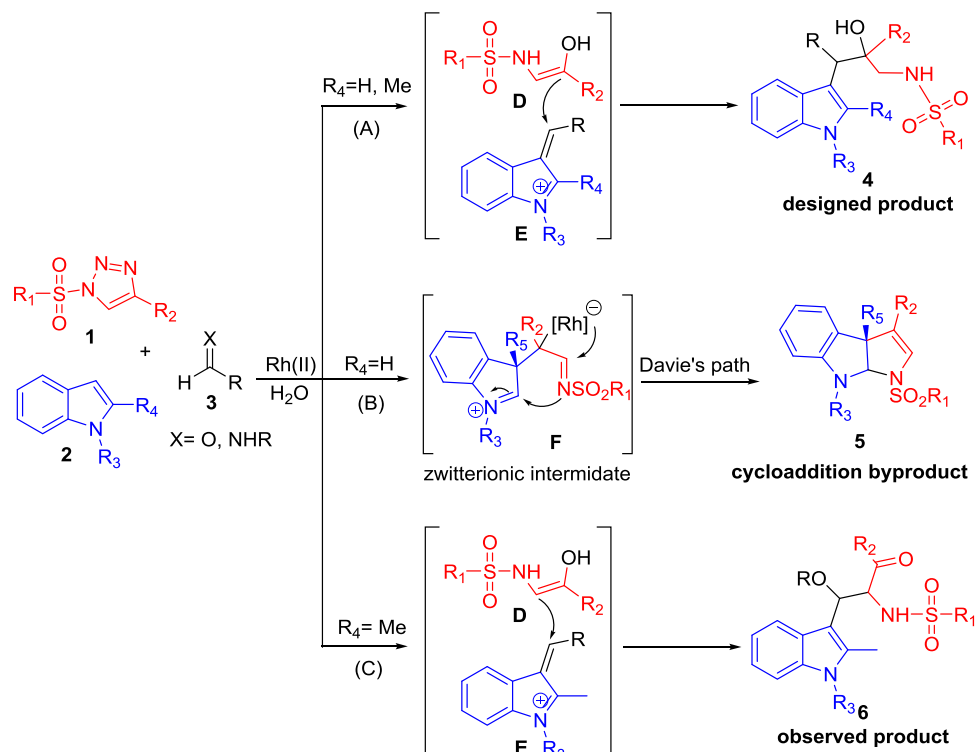
The reaction was then conducted with 4-(4-bromophenyl)-1-(methylsulfonyl)-1H-1,2,3-triazole **1a** (1.0 eq), an excess of 1,3-dimethyl indole **2** (which could possibly generate the zwitterionic intermediate **F**, as suggested by Davies), and electrophile **3** (4.0 eq), such as an imine, isatin, aromatic aldehyde, ethyl glyoxylate, formaldehyde solution, and paraformaldehyde, in the presence of 2.0 mole percent (mol %) of Rh<sub>2</sub>(OAc)<sub>4</sub>. Disappointingly, no multicomponent product was observed even when the N-CH<sub>3</sub> indole was used instead of the 1,3-dimethyl indole, and the major product was compound **5**, as illustrated by Davies *et al.* In view of the charge effect of the electron-donating substituents on the nucleophilicity of the C(3) on the indole ring, 1,2-dimethyl indole **2** was then used instead of 1,3-dimethyl indole to investigate the possible reaction outcome (Table 1). Some multicomponent products were detected by high-performance liquid chromatography–mass spectrometry (HPLC-MS) and <sup>1</sup>H nuclear magnetic resonance (NMR) spectroscopic analyses when ethyl glyoxylate was used as trapping reagent, but it was not possible to obtain pure products because of their low yield. Gratifyingly, an aqueous solution of formaldehyde afforded a trapping product with the desired molecular weight of 439 in 20% isolated yield. Encouraged by these positive results, paraformaldehyde was used without additional water to explore the MCR and a better yield of 29% was obtained. Both formaldehyde aqueous solution and paraformaldehyde provided major by-product **5** in a ~28% isolated yield. When using 1,3-dimethyl indole, other common electrophiles, such as imines, isatins, or aromatic aldehydes, also provided no desired three-component product but two-component by-product **5**.

### Optimization of the MCR conditions

To evaluate the efficiency of this new transformation, we examined the reaction conditions using paraformaldehyde, and the optimal

conditions are illustrated in Table 2. As shown,  $\text{Rh}_2(\text{OAc})_4$  was more effective than  $\text{Rh}_2(\text{Oct})_4$ ,  $\text{Rh}_2(\text{Piv})_4$ ,  $\text{Rh}_2(\text{TPA})_4$  and  $\text{Rh}_2(\text{S-NTTL})_4$  in providing the desired product **6a** in 29% yield in toluene at 100°C (entries 1 to 5). Chlorobenzene was a better solvent than toluene, ethyl acetate (EA), cyclohexane, tetrahydrofuran (THF), dimethylbenzene,

$\text{CHCl}_3$ , and 1,2-dichloroethane (DCE), leading to a higher 36% yield (entries 6 to 12). No desired product was obtained from cyclohexane (entry 9), possibly because of the fact that a nonpolar hydrocarbon solvent might be unfavorable for the generation of the vinylimine ion intermediate. It was found that the reaction temperature had a profound

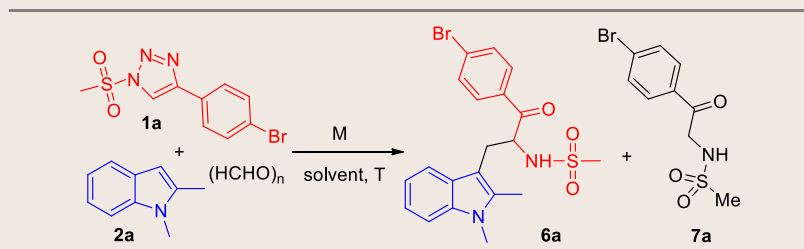


**Scheme 2. Designed MCR pathway.** Designed and observed transformation processes of trapping of the  $\alpha$ -amino enol **D** with indole vinylimine ion **E** (**A** and **C**) versus competitive Davies' cycloaddition pathway (**B**).

**Table 1. Preliminary study of the MCR.** The preliminary study of the MCR by trapping  $\alpha$ -amino enol **D** with various electrophiles. Unless otherwise stated, all reactions were carried out in PhMe in a 0.1-mmol scale and **1a:2a:3** = 1.0:5.0:4.0 and 2 mol % of  $\text{Rh}_2(\text{OAc})_4$  catalyst were added. Yields of isolated products were obtained after purification by column chromatography. N.P., no desired product.

Imine	Isatin	Aromatic aldehyde	Ethyl glyoxylate	Formaldehyde	Paraformaldehyde
Yield:	N.P.	N.P.	~8% yield R = CO <sub>2</sub> Et	20% yield R = H	29% yield R = H

**Table 2. Reaction condition screening.** The optimization studies of the MCR of trapping  $\alpha$ -amino enol **D** with paraformaldehyde. Unless otherwise stated, all reactions were carried out in solvents in a 0.1-mmol scale and **1a:2a:3** = 1.0:5.0:4.0 and 2 mol % of metal catalyst were added. Yields of isolated products were obtained after purification by column chromatography.



Entry	[M]	Solvent	T(°C)	Yield (%)
1	Rh <sub>2</sub> (Oct) <sub>4</sub>	PhMe	100	27
2	Rh <sub>2</sub> (Piv) <sub>4</sub>	PhMe	100	18
3	Rh <sub>2</sub> (TPA) <sub>4</sub>	PhMe	100	24
4	Rh <sub>2</sub> (S-NTTL) <sub>4</sub>	PhMe	100	22
5	Rh <sub>2</sub> (OAc) <sub>4</sub>	PhMe	100	29
6	Rh <sub>2</sub> (OAc) <sub>4</sub>	CHCl <sub>3</sub>	100	21
7	Rh <sub>2</sub> (OAc) <sub>4</sub>	Dimethylbenzene	100	25
8	Rh <sub>2</sub> (OAc) <sub>4</sub>	EA	100	11
9	Rh <sub>2</sub> (OAc) <sub>4</sub>	Cyclohexane	100	0
10	Rh <sub>2</sub> (OAc) <sub>4</sub>	THF	100	19
11	Rh <sub>2</sub> (OAc) <sub>4</sub>	1,2-DCE	100	31
12	Rh <sub>2</sub> (OAc) <sub>4</sub>	PhCl	100	36
13	Rh <sub>2</sub> (OAc) <sub>4</sub>	PhCl	60	28
14	Rh <sub>2</sub> (OAc) <sub>4</sub>	PhCl	80	32
15	Rh <sub>2</sub> (OAc) <sub>4</sub>	PhCl	120	49
<b>16</b>	<b>Rh<sub>2</sub>(OAc)<sub>4</sub></b>	<b>PhCl</b>	<b>140</b>	<b>70</b>
17	Rh <sub>2</sub> (OAc) <sub>4</sub>	PhCl	160	43
18	Rh <sub>2</sub> (S-NTTL) <sub>4</sub>	PhCl	140	29
19	Rh <sub>2</sub> (Oct) <sub>4</sub>	PhCl	140	21
20	Rh <sub>2</sub> (TPA) <sub>4</sub>	PhCl	140	16

influence on the yield, and an optimal reaction temperature of 140°C was identified and effectively improved the yield to 70% (entries 12 to 17). A very little amount of another by-product, ketone **7a**, which was generated from  $\alpha$ -amino enol **D** by keto-enol tautomerization, was observed when the reaction temperature was increased to 140°C. Other metal catalysts still produced a much poor yield at the optimal reaction temperature of 140°C in chlorobenzene (entries 18 to 20). At the reaction temperature of 140°C, chlorobenzene was also demonstrated as the better solvent than other solvents when using Rh<sub>2</sub>(OAc)<sub>4</sub> as catalyst (see table S1). Under the optimized conditions, the yield of by-product **5** was obviously suppressed to ~10%. The MCRs of 1-sulfonyl-1,2,3-triazoles as precursors of Rh(II) azavinyl carbenes are really restricted, although the involved two-component reactions are very fruitful (19). Here, we envisioned that a Rh(II)-catalyzed MCR with 1-sulfonyl-1,2,3-triazoles as precursors of azavinyl carbenes could be successfully developed.

#### Determination of the MCR products

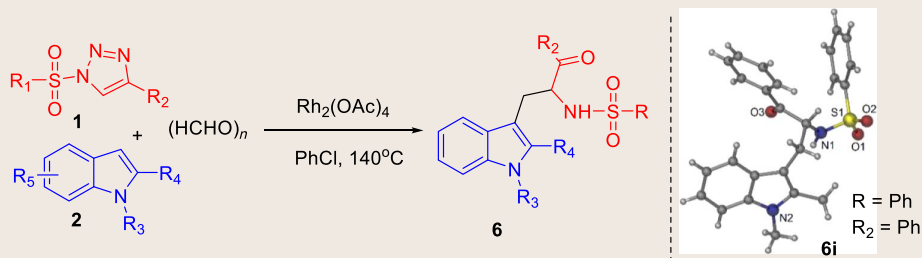
All the products were respectively characterized by <sup>1</sup>H NMR, <sup>13</sup>C NMR, and high-resolution mass spectrometry (HRMS) (see the Supplementary Materials). Surprisingly, the structure of the multicomponent prod-

uct corresponded to  $\alpha$ -amino- $\beta$ -indole ketone **6** rather than the designed product **4**, as determined using single-crystal x-ray analysis by analogy with **6i** (see figure in Table 3 and fig. S2). This suggests a selectivity of the nucleophilicity between  $\alpha$ -C and  $\beta$ -C in enol **D** during the trapping process.

#### Substrate scope of the MCR

The scope of this effective process was then surveyed with respect to the triazole and indole coupling partners. As shown in Table 3, a broad range of triazoles **1** provided the corresponding products in moderate to good yields (entries 1 to 9). 1-MesyI-substituted triazoles bearing a bromide or other electron-rich groups at the C-4' of the phenyl ring produced good yields of the corresponding products (entries 1 to 3). 1-MesyI-substituted triazole **1d** featuring a heterocyclic group also produced a good yield of the desired product (entry 4). By-product **7c** was also observed and isolated in 10% yield when an OEt (R<sub>2</sub>) substituent was present at the C-4' on the phenyl ring (entry 3), whereas electron-deficient groups (for example, -CF<sub>3</sub> and -COOMe) provided trace amounts of the desired product (entry 5). Bulky triazoles **1** were found to be tolerated despite the steric effect and can equally furnish the

**Table 3. Substrate scope study.** The substrate scope of the MCR by using various triazole and indole components. Unless otherwise stated, all reactions were carried out in PhCl in a 0.2-mmol scale at 140°C and 1:2:3 = 1.0:5.0:4.0 and 2 mol % of Rh<sub>2</sub>(OAc)<sub>4</sub> were added.



Entry	<b>6</b>	R <sub>1</sub>	R <sub>2</sub>	R <sub>3</sub>	R <sub>4</sub> /R <sub>5</sub>	Yield (%) <sup>*</sup>
1	<b>6a</b>	Me	4-BrPh	Me	Me/H	70
2	<b>6b</b>	Me	4-OMePh	Me	Me/H	55
3	<b>6c</b>	Me	4-OEtPh	Me	Me/H	60
4	<b>6d</b>	Me	3-Thienyl	Me	Me/H	75
5	<b>6e</b>	Me	4-CF <sub>3</sub> Ph	Me	Me/H	NP <sup>†</sup>
6	<b>6f</b>	<i>i</i> -Pr	4-MePh	Me	Me/H	45
7	<b>6g</b>	4-BrPh	4-MePh	Me	Me/H	50
8	<b>6h</b>	4-OCF <sub>3</sub> Ph	4-MePh	Me	Me/H	64
9	<b>6i</b>	Ph	Ph	Me	Me/H	71
10	<b>6j</b>	4-OMePh	4-MePh	Me	Me/H	33
11	<b>6k</b>	Me	4-BrPh	H	Me/H	55
12	<b>6l</b>	Me	4-BrPh	Bn	Me/H	62
13	<b>6m</b>	Me	4-BrPh	Me	Ph/H	52
14	<b>6n</b>	Me	4-BrPh	Me	CO <sub>2</sub> Et/H	45
15	<b>6o</b>	Me	4-BrPh	Me	Me/Me	51
16	<b>6p</b>	Me	4-BrPh	Me	Me/OMe	<10

<sup>\*</sup>Isolated yields after purification by column chromatography. <sup>†</sup>NP: A trace amount of product was monitored by LC-MS.

corresponding products (entries 6 to 10). This transformation was also applicable to a range of C(2)-substituted indoles **2** (Table 3, entries 11 to 15). An evaluation of the substituents on the indolic nitrogen revealed that N-H and N-Bn indoles were both compatible with the reaction conditions to provide the products in 55 and 62% yield, respectively (entries 11 and 12). It is unclear why a methoxy group at the C(5) of the indole core produced a low yield of 10% (entry 16).

### Synthetic application of the MCR

The synthetic application of multicomponent product **12** was illustrated by the simple reduction of ketone **6c** to alcohol **11**, followed by a ring closure with triphosgene to afford 3-(methylsulfonyl)oxazolidin-2-one derivative **12**, which frequently serves as a pivotal pharmaceutical moiety (Scheme 3) (71–73). The relative configuration of **12** was determined to be (1*S*, 2*R*) on the basis of the corresponding two-dimensional (2D) <sup>1</sup>H-<sup>1</sup>H nuclear Overhauser effect spectroscopy (NOESY) spectrum (see fig. S3).

### Experimental mechanistic investigation

This effective chemical transformation strongly motivated us to explore its reaction mechanism. The observation of the formation of by-product **7** indicated that intermediate enol **D** generated from water and carbene **A** might be involved in this MCR. The reaction was thus performed with the original three components and heavy oxygen water (H<sub>2</sub>O<sup>18</sup>)

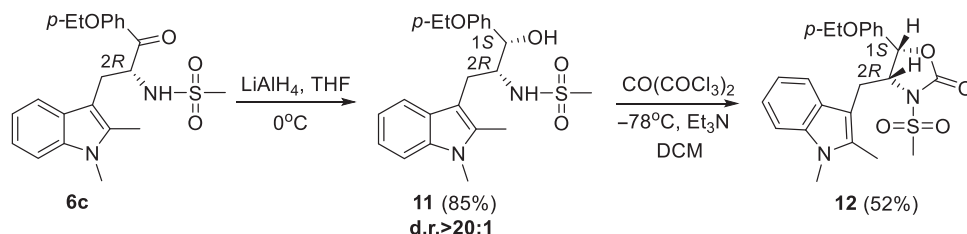
to provide the isotope-labeled product **6i'** (Scheme 4A). The molecular ion peak of **6i'** was observed at 432.2 and 434.2 (calculated [M + H]<sup>+</sup> for **6h'** is 433.2 and 435.2) (see fig. S4). However, when 4 Å molecular sieves were added to the reaction solution, the desired product did not form (Scheme 4B). The results of these control experiments confirmed that water participated in the multicomponent process, and the reaction could be run in air. A sample of the reaction mixture solution was analyzed by HPLC-MS to monitor the progress of the reaction, and the results showed that a molecular ion peak was observed at 174.9, which might be attributed to vinylimine ion **E** generated from indole and HCHO (see fig. S5; calculated [M + H]<sup>+</sup> for **E** was 174.9).

Considering that triazoles can be decomposed by heating, the reaction was conducted at 140°C in chlorobenzene without Rh<sub>2</sub>(OAc)<sub>4</sub>, and no desired product was observed (Scheme 4C), which indicated that the reaction was catalyzed by Rh<sub>2</sub>(OAc)<sub>4</sub>. A further control experiment was conducted by reacting triazole **1i** with paraformaldehyde under the identified conditions to exclude another possible pathway, according to which the HCHO molecules would first react with α-amino enol **D** to form **8** followed by an electrophilic attack onto the indole. No two-component product **8** but ketone **7** and a little amount of [3 + 2] cycloaddition product **9** were observed (Scheme 4D). These results imply that the HCHO molecule first reacted with C(2)-substituted indole **2** to generate an intermediate, which subsequently reacted with α-amino enol **D**.

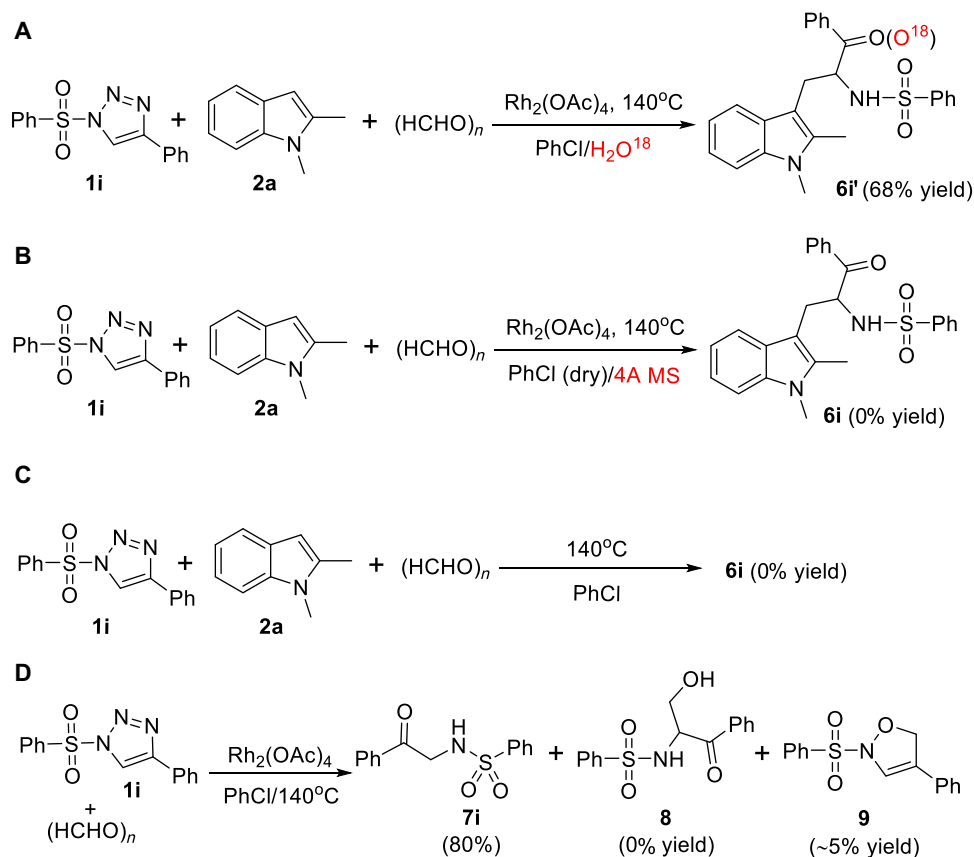
### Mechanistic insight based on a theoretical calculation investigation

To further understand the two competitively crucial transformation pathways of  $\alpha$ -amino enol **D** into ketone **7** by keto-enol tautomerization versus trapping by vinylimine intermediate **E** into product **6**, we performed detailed DFT calculations for the transformation of the two key intermediates, for example,  $\alpha$ -amino enol **D** and vinylimine ion **E**. The typically calculated free energy profiles for the formation and the elementary transformations of  $\alpha$ -amino enol **D** are shown in Fig. 2. First, we were interested in how to generate  $\alpha$ -amino enol **D**. It is well known that Rh(II)-associated carbenes readily react with water/alcohol to form the corresponding active oxonium ylides, which can yield O–H insertion products through a proton transfer process (74–76). Xia *et al.* (77) also demonstrated this process by theoretical calculations. Our group further demonstrated that these oxonium ylides could be

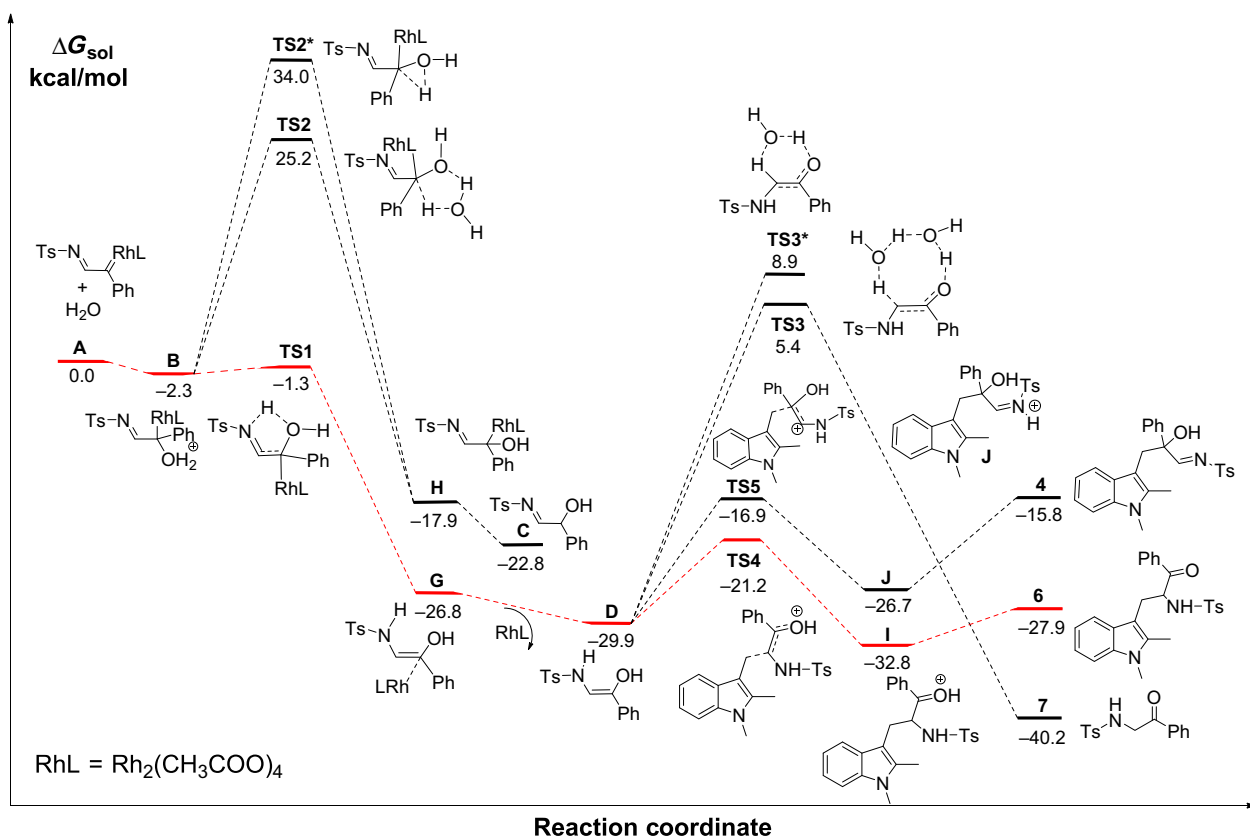
trapped by electrophiles during the experiments before the proton transfer took place (78, 79). Furthermore, both free energy profiles for active ylide intermediate **B** (52) and water O–H bond insertion intermediate **C** were optimized for the transformation of  $\alpha$ -imino Rh(II) carbene **A**, as proposed by Miura *et al.* (51). It was found that the formation of ylide **B** from reactant **A** and one water molecule had no barrier. The transformation of ylide **B** to enol **D** was extremely exothermic by 29.9 kcal/mol via a specific five-membered ring transition state **TS1**, suggesting that intermediate **D** is easily formed via ylide **B**. Ylide **B** can also easily transfer to intermediate **C** via a water-associated proton shift according to the calculations by Liang *et al.* (80). In our calculated energy profiles, the activation energy barrier for the transformation of ylide **B** into intermediate **C** was 36.3 kcal/mol via a concerted transition state **TS2\*** and 27.5 kcal/mol via a water-associated **TS2**, which indicates that this transformation pathway is unfavorable. Then, the proposed key process,



**Scheme 3. Synthetic application of the products.** The synthetic application of the multicomponent products for the rapid construction of 3-(methylsulfonyl)oxazolidin-2-one derivatives. d.r., diastereomeric ratio; DCM, dichloromethane.



**Scheme 4. Control experiments.** Control experiments to obtain an insight into the MCR pathway: **(A)** and **(B)** The results indicate that water is involved in the MCR; **(C)** the result indicates that  $\text{Rh}_2(\text{OAc})_4$  rather than heating promoted the reaction; and **(D)** the results exclude the other possible pathway according to which paraformaldehyde would first react with triazole.

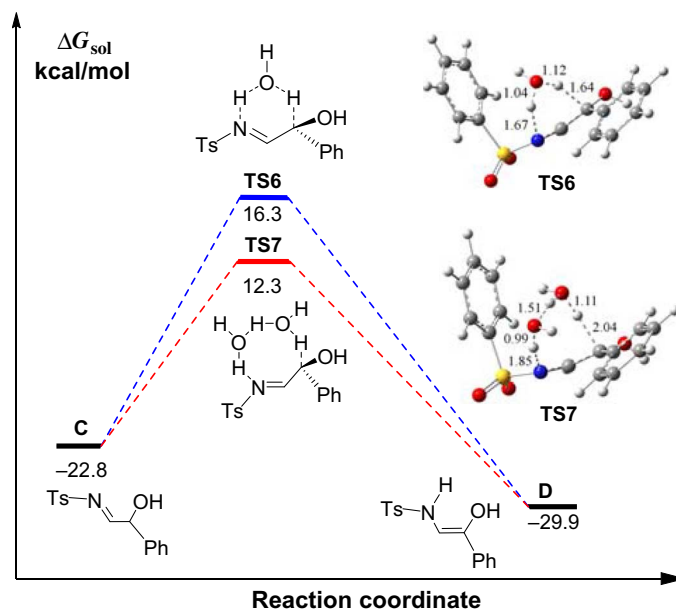


**Fig. 2. DFT investigation of the MCR.** Calculated free energy profiles at 140°C in chlorobenzene for the generation and transformation of enol **D** into ketones and into the multicomponent product with vinylimine intermediate **E**. All the energy values are expressed in kilocalorie per mole.

namely, the imine-enamine tautomerization from intermediate **C** to enol **D**, was investigated. Considering that proton shuttles might greatly facilitate the imine tautomerization process to enamine, water molecules were included in the transformation. The calculated activation energy barrier for the proton transfer in the imine-enamine tautomerization process from intermediate **C** to enol **D** was high, up to 39.1 kcal/mol via **TS6** assisted by one water molecule and up to 35.1 kcal/mol via **TS7** assisted by two water molecules (Fig. 3), which implies that the conversion from intermediate **C** to **D** is impossible.

Subsequently, the key transition states of the transformation of enol **D** into ketone **7** and product **6** were located in the energy profiles (Fig. 2), respectively. The calculated activation energy barrier for **TS3** in which two water molecules play as proton shuttles (80) for the transformation of enol **D** into ketone **7** was 35.3 kcal/mol at 140°C. The energy barrier through the one-water-associated **TS3\*** was evaluated as 38.8 kcal/mol, which suggests a more unfavorable pathway in terms of kinetics. In general, the keto-enol tautomerization in solution is an equilibrium reaction. The calculated reverse reaction barrier from ketone **7** to enol **D** was as high as 45.6 kcal/mol, which rules out the possibility of a reverse transformation from ketone **7** to enol **D**. The theoretically calculated results are in good agreement with the observations in our control experiments that multicomponent product **6** could not be obtained by reacting the isolated ketone **7** with indole **2** and paraformaldehyde under the same reaction conditions.

The key transition state **TS4** for the nucleophilic addition of enol **D** to vinylimine intermediate **E** in chlorobenzene at 140°C was also obtained. As shown in Fig. 2, the evaluated barrier for the forward reaction

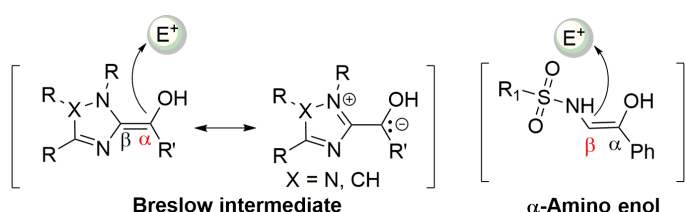


**Fig. 3. DFT investigation of the imine-enamine tautomerization.** Calculated free energy profiles for the imine-enamine tautomerization from intermediate **C** to enol **D**. The blue line denotes the pathway via the transition state **TS6** through one water molecule, whereas the red curve denotes the pathway via **TS7** through two water molecules, respectively.

was 8.7 kcal/mol, appearing more favorable compared to the 35.3 kcal/mol of **TS3** in the keto-enol tautomerization, which illustrates the dominant transformation of enol **D** with vinyimine intermediate **E** to generate product **6**. The calculated barrier for this step is also consistent with the experimental observation that the multicomponent transformation was rapidly completed in half an hour. The DFT calculations on the reaction mechanism suggest that the active  $\alpha$ -amino enol **D** was trapped by the in situ-generated vinyimine ion **E** to promote the MCR by an intermolecular reassembly.

To explain the selectivity of the nucleophilicity between  $\alpha$ -C and  $\beta$ -C in enol **D**, the activation energy barrier of the forward reaction at the  $\alpha$ -C site of the enol attacked by vinyimine intermediate **E** via **TS5** was also evaluated (Fig. 2), and a value of 13.0 kcal/mol was obtained. The barrier of **TS5** was higher than that of **TS4** by 4.3 kcal/mol, suggesting a lower probability for this precursor to react with intermediate **E** at the  $\alpha$ -C in enol **D** to yield product **4**. Moreover, the free energy of product **4** was higher than that of enol **D** by 12.1 kcal/mol, indicating that the forward reaction from enol **D** to product **4** is unfavorable from a thermodynamic point of view.

It is interesting and noteworthy that the water molecules play a crucially important role in the formation of vinyimine ion **E** to promote the multicomponent process on the basis of the DFT calculations (see fig. S6). The theoretical results support the assumption that a hydrated proton, but not a free proton, promoted the generation of vinyimine



**Scheme 5. Selective nucleophilic attack of the intermediate.** Different nucleophilic attack selectivity of  $\alpha$ -amino enol **D** versus Breslow intermediate.

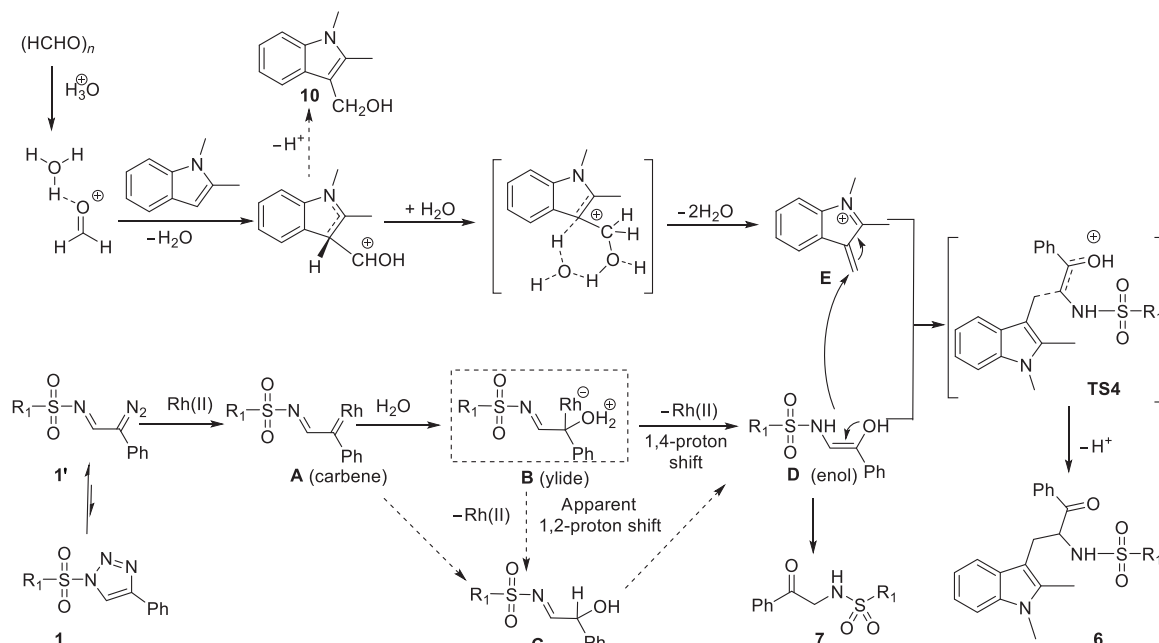
ion **E**, which is consistent with our control experimental results that a Brønsted acid, such as phosphoric acid, could not accelerate or improve the multicomponent transformation. Possibly because of the difference between the electronic features of paraformaldehyde and alkyl or aryl aldehyde (Scheme 2), no intermediate product **10** could be located in the DFT energy profile (fig. S6), which is also in accordance with the observation that no indole product **10** was obtained experimentally.

## DISCUSSION

The determination of the product structure indicates that a selective nucleophilic attack occurred at the enolic carbon ( $\beta$ -C) rather than at the enamic carbon ( $\alpha$ -C) of the key amino enol intermediate **D** (Scheme 2C versus Scheme 2A). Interestingly, to some extent, the structure of  $\alpha$ -amino enol intermediate **D** is similar to that of N-heterocyclic carbene (NHC) intermediates (81–86), such as the Breslow intermediate in the Stetter reaction (Scheme 5) (87, 88). Breslow intermediates are classical examples of umpolung chemistry in which the  $\alpha$ -C (enamine reactivity) rather than the  $\beta$ -C (enol reactivity) dominates the nucleophilic attack (89–91). Although both  $\alpha$ -C and  $\beta$ -C in  $\alpha$ -amino enol **D** can theoretically perform a nucleophilic addition to vinyimine intermediate **E**, the enolic carbon ( $\beta$ -C) here unexpectedly exhibited a higher reactivity than that of the enamic carbon, leading to the formation of  $\alpha$ -amino- $\beta$ -indole ketones.

The substrate scope survey demonstrates the efficiency of this MCR in affording  $\alpha$ -amino- $\beta$ -indole ketones, which exhibit important bio-activity profiles in drug discovery (92, 93) and can be used as synthetic intermediates for the production of  $\alpha$ -amino alcohols (94–97) or natural products (98, 99). To the best of our knowledge, this represents the first example of an effective multicomponent sequence that rapidly affords  $\alpha$ -amino- $\beta$ -indole ketones.

For the classical two-component transformation of  $\alpha$ -imino Rh(II) carbene **A** (which can be generated in situ from the reaction of 1-sulfonyl-1,2,3-triazole with water) into the corresponding ketone, a



**Scheme 6. Plausible MCR mechanism.** Proposed MCR mechanism of the  $\alpha$ -amino enol **D** trapping by indole vinyimine ion **E**.



mechanism was proposed (51) via key intermediate  $\alpha$ -amino enol **D** by imine-enamine tautomerization from water O–H bond insertion intermediate **C**, subsequently followed by keto-enol tautomerization via a proton transfer process. Our investigations based on detailed DFT calculations and control experiments have excluded possible transformations of both imine-enamine tautomerization from **C** to **D** and keto-enol tautomerization from ketones (products) to  $\alpha$ -amino enol **D**. An extremely favored energy profile suggests that  $\alpha$ -imino Rh(II) carbene **A** can possibly yield key intermediate  $\alpha$ -amino enol **D** via metal-associated ylide **B**.

On the basis of the abovementioned results, a plausible reaction pathway was proposed, as shown in Scheme 6. Initially, a ring-chain tautomerization of *N*-sulfonyl-1,2,3-triazole **1** generates  $\alpha$ -diazo imine **1'** (100–102). Diazo tautomer **1'** is decomposed by a rhodium(II) complex with release of N<sub>2</sub> to provide  $\alpha$ -imino rhodium carbene **A**, which reacts with water to generate Rh(II)-associated ylide **B**. This active ylide **B** subsequently undergoes a 1,4-proton transfer to afford key intermediate  $\alpha$ -imino enol **D**, regenerating the rhodium(II) catalyst. Simultaneously, the other key intermediate, vinylimine intermediate **E**, is generated in situ from indole and paraformaldehyde via a water-associated transition state. Finally, vinylimine intermediate **E** is attacked by  $\alpha$ -amino enol **D** to generate the multicomponent product with release of water.

In summary, an MCR based on the trapping of an active  $\alpha$ -imino enol intermediate in a traditional two-component reaction pathway was successfully illustrated. This transformation was realized by trapping active  $\alpha$ -imino enol intermediates obtained using 1-sulfonyl-1,2,3-triazoles as precursors by vinylimine ions generated in situ from indoles and paraformaldehyde to suspend the traditional enol-ketone transformation pathway in favor of a new multicomponent pathway in the presence of rhodium(II) acetate as catalyst. Compared to classical NHC intermediates, the enolic carbon had a stronger nucleophilicity than that of the enamic carbon in the trapping process. Both series of control experiments and detailed DFT theoretical calculations aided to demonstrate the MCR mechanism. The developed MCR provides a mild, efficient, and convenient access to  $\alpha$ -amino- $\beta$ -indole ketones, and further synthetic application of the resulting products was illustrated.

Although the presently developed MCR has a substrate scope limitation, its extension to other electrophiles other than paraformaldehyde as trapping reagents, such as benzaldehyde, could provide a mean to achieve an asymmetric version of this MCR. Further efforts on extending this process to different reactants and trapping other active intermediates for the discovery of new MCRs are currently being pursued in our laboratory.

## MATERIALS AND METHODS

### Experimental design

All <sup>1</sup>H NMR (400 MHz) and <sup>13</sup>C NMR (100 MHz) spectra were recorded on Bruker spectrometers using CDCl<sub>3</sub> as solvent. Tetramethylsilane served as internal standard ( $\delta = 0$ ) for <sup>1</sup>H NMR, whereas CDCl<sub>3</sub> was used as internal standard ( $\delta = 77.0$ ) for <sup>13</sup>C NMR. Chemical shifts were reported in parts per million with the following multiplicity: singlet, doublet, triplet, quartet, multiplet, and broad. HRMS was performed on IonSpec Fourier transform-ion cyclotron resonance or Waters Micromass Q-TOF micro synapt high definition mass spectrometry. Melting points were uncorrected. Single-crystal x-ray diffraction data (**6i**) were recorded on a Bruker-AXS SMART APEX II single-crystal x-ray diffractometer.

The solvents used for column chromatography were distilled before use. All reactions and manipulations were carried out in air using a flame- or oven-dried flask containing a magnetic stirring bar. 1-Sulfonyl-substituted triazoles and indoles were prepared according to the literature (103, 104).

### General procedure for the synthesis of 1-sulfonyl-substituted triazoles 1

A solution of sodium azide (4.88 g, 75 mmol) in 10 ml of water at 0°C was added slowly to a stirred solution of methanesulfonyl chloride (5.73 g, 50 mmol) in 90 ml of acetone. The reaction mixture was stirred at this temperature for 1 hour and then for 2 hours at room temperature. The solution was concentrated to about 25 ml, and it was then extracted with diethyl ether (2 × 50 ml), washed with 10 ml of water and then brine, and dried over anhydrous magnesium sulfate. After filtration, the resulting filtrate containing the crude mesyl azide was concentrated to about 10 ml and used in the next step without further purification. Copper(I) thiophene-2-carboxylate (0.190 g, 1.0 mmol, 2 mol %) was added to a stirred solution of phenylacetylene (5.1 g, 50 mmol) in toluene (100 ml) at room temperature. After stirring for 1 to 3 min, a solution of the crude methanesulfonyl azide in diethyl ether was added dropwise to the resulting suspension. The reaction mixture was stirred at room temperature for about 12 hours until completion, as confirmed by liquid chromatography mass spectrometry (LC-MS) or thin-layer chromatography (TLC) analysis, and it was then concentrated under reduced pressure and filtered through a short plug of silica to remove the copper catalyst (40% EtOAc/hexane as eluent). After removal of all the solvents in vacuum, the remaining off-white solid was triturated with EtOAc (15 ml) and hexane (70 ml) to afford 7.2 g (32.5 mmol, 65% over two steps) of 1-(methylsulfonyl)-4-phenyl-1*H*-1,2,3-triazole as a white crystalline solid.

### General procedure for establishing the preliminary reaction conditions (Table 1)

Triazole **1** (0.1 mmol, 1.0 eq), indole **2** (0.5 mmol, 5.0 eq), electrophile **3** (0.4 mmol, 4.0 eq), and Rh catalyst (0.02 mmol, 0.02 eq) were added to an oven-dried test tube. The test tube was placed under an air atmosphere. Undistilled toluene (2.0 ml) was added, and the reaction was heated at 90°C under vigorous stirring. After completion of the reaction (monitored by TLC or LC-MS), the mixture was cooled to room temperature and all the volatiles were removed in vacuum. The crude residue was purified by flash chromatography [EA/petroleum ether (PE), 1:15 to 1:8] to provide the title compound.

### General procedure for the optimization of the reaction conditions (Table 2)

Triazole **1** (0.1 mmol, 1.0 eq), indole **2** (0.5 mmol, 5.0 eq), paraformaldehyde **3** (0.4 mmol, 4.0 eq), and Rh catalyst (0.02 mmol, 0.02 eq) were added to an oven-dried test tube. The test tube was placed under an air atmosphere. Undistilled solvent (2.0 ml) was added, and the reaction was heated at the indicated temperature under vigorous stirring. After completion of the reaction (monitored by TLC or LC-MS), the mixture was cooled to room temperature and all the volatiles were removed in vacuum. The crude residue was purified by flash chromatography (EA/PE, 1:15 to 1:8) to provide the title compound.

### General procedure for the synthesis of $\alpha$ -amino- $\beta$ -indole ketones (Table 3)

Triazole **1** (0.2 mmol, 1.0 eq), indole **2** (1.0 mmol, 5.0 eq), paraformaldehyde **3** (0.8 mmol, 4.0 eq), and Rh<sub>2</sub>(OAc)<sub>4</sub> (1.7 mg, 0.04 mmol,

0.02 eq) were added to an oven-dried test tube. The test tube was placed under an air atmosphere. Chlorobenzene (2.0 ml) was added, and the reaction was heated at 140°C under vigorous stirring for 15 to 30 min. After completion of the reaction (monitored by TLC or LC-MS), the mixture was cooled to room temperature and all the volatiles were removed in vacuum. The resulting crude residue was purified by flash chromatography (EA/PE, 1:15 to 1:8) to provide the title compound.

### General procedure for the control experiment

Indole **2** (0.5 mmol, 5.0 eq), triazole **1** (0.1 mmol, 1.0 eq), paraformaldehyde (0.4 mmol, 4.0 eq), H<sub>2</sub>O<sup>18</sup> (1.0 mmol, 10.0 eq), and Rh<sub>2</sub>(OAc)<sub>4</sub> (0.9 mg; 0.02 mmol, 0.02 eq) were added to an oven-dried test tube. The test tube was placed under an air atmosphere. Chlorobenzene (1.0 ml) was added, and the reaction was heated at 140°C under vigorous stirring for 15 to 30 min. After completion of the reaction (monitored by TLC or LC-MS), the mixture was cooled to room temperature and all the volatiles were removed in vacuum. The crude residue was purified by flash chromatography (EA/PE, 1:15 to 1:8) to provide the title compounds **6i** and **6i'**.

Triazole **1i** (0.1 mmol, 1.0 eq), paraformaldehyde (0.4 mmol, 4.0 eq), 4 Å molecular sieves (30 mg), and Rh<sub>2</sub>(OAc)<sub>4</sub> (0.9 mg; 0.02 mmol, 0.02 eq) were added to an oven-dried test tube. The test tube was placed under an air atmosphere. Chlorobenzene (1.0 ml) was added, and the reaction was heated at 140°C under vigorous stirring for 15 to 30 min. After completion of the reaction (monitored by TLC), the mixture was cooled to room temperature and all the volatiles were removed in vacuum. The reaction was monitored by LC-MS, but we failed to obtain the two-component product **8** of the reaction of **1i** with paraformaldehyde.

### General procedure for the synthesis of oxazolidin-2-one **12** (**105**)

A 2.5 M LiAlH<sub>4</sub> solution (0.4 mmol) in THF was added to a stirred solution of **6c** (82 mg, 0.2 mmol) in THF (2.0 ml) at 0°C. The mixture was stirred at 0°C for 1 hour. After completion of the reaction (monitored by TLC or LC-MS), the mixture was quenched with 1 M aq. HCl and extracted with EA (×2). The combined organic layers were washed with saturated aq. NaHCO<sub>3</sub> and brine and then dried over MgSO<sub>4</sub>. After evaporation under reduced pressure, the residue was purified by silica gel flash column chromatography (PE/EA, 10:1 to 3:1) to provide α-amino alcohol **11** (71 mg, yield 85%) as a yellow solid. Pyridine (0.1 ml) and triphosgene (19 mg, 0.06 mmol) in CH<sub>2</sub>Cl<sub>2</sub> (0.5 ml) were added to a stirred solution of **11** (20 mg, 0.05 mmol) in CH<sub>2</sub>Cl<sub>2</sub> (1.0 ml) at -78°C. The mixture was stirred at -78°C for 20 min and then was gradually warmed to room temperature for more than 1 hour. The mixture was quenched with 1 M aq. HCl and extracted with EA (×2). The combined organic layers were washed with saturated aq. NaHCO<sub>3</sub> and brine and then dried over MgSO<sub>4</sub>. After evaporation under reduced pressure, the residue was purified by silica gel flash column chromatography (PE/EA, 10:1 to 3:1) to provide oxazolidin-2-one **12** (10 mg, yield 52%) as a white solid.

### Computational methods

All the DFT calculations were carried out by using the Gaussian09 software package (**106**). The structures of intermediates and transition states were optimized using the M06 (**107**) functional combined with the 6-31G\* (**108**) basis set for nonmetal elements as well as the LanL2dz (**109**) basis set for metal elements, which is denoted as the M06/LanL2dz +6-31G\* method. The single-point energy calculations and frequency analyses were further performed on the basis of the optimized

structures, with the larger basis sets of 6-311G\*\* (**108**) basis set used for nonmetal elements and the Stuttgart/Dresden (SDD) (**110**) basis set used for metal elements, denoted as the M06/SDD+6-311G\*\* method. The solution effect of chlorobenzene was estimated using the Solvation Model based on Density (SMD) model (**111**) in Gaussian09 on the basis of the optimized structures in gas phase. All the energies presented in the content refer to the Gibbs free energies ΔG<sub>sol</sub> calculated at the temperature of 140°C, including the corrections of solvation energies in chlorobenzene.

### SUPPLEMENTARY MATERIALS

Supplementary material for this article is available at <http://advances.sciencemag.org/cgi/content/full/3/3/e1602467/DC1>

fig. S1. Rate-determining step of the proton transfer process compared with the active intermediate generation step in the reaction coordinate of the X-H (X = O, N) insertion product formation.

fig. S2. Single-crystal x-ray structure determination of compound **6i** (CCDC no. 1407544).

fig. S3. Two-dimensional <sup>1</sup>H-<sup>1</sup>H NOESY spectrum of compound **12**.

fig. S4. LC-MS spectra of the reaction solution containing indole, triazole, paraformaldehyde, H<sub>2</sub>O<sup>18</sup>, and Rh<sub>2</sub>(OAc)<sub>4</sub>.

fig. S5. LC-MS spectra of the reaction solution containing indole, triazole, paraformaldehyde, and Rh<sub>2</sub>(OAc)<sub>4</sub>.

fig. S6. Calculated free energy profiles of the formation of vinylimine ion **E** at 140°C in chlorobenzene.

table S1. Reaction condition screening.

data file S1. Characterization of products.

data file S2. X-ray crystal data.

data file S3. NMR spectra of compounds.

data file S4. DFT calculation data.

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**Acknowledgments:** We thank X. Zhao for helping with the single-crystal x-ray analysis.

**Funding:** We thank the National Science of Foundation of China (21332003, 21473056, and 21672066), the Science and Technology Commission of Shanghai Municipality (15ZR1411000), and the State Key Laboratory of Advanced Technology for Materials Synthesis and Processing (2015-KF-11) for financial support. **Author contributions:** S.L. designed the experimental protocols and wrote the paper. W.Y. executed the main experimental work. Q.W. carried out the control experiments. F.X., Y.L., and J.C. performed the DFT calculation investigation. X.W. contributed to the determination of the product structures. W.H. proposed the idea and designed the investigation proposal. **Competing interests:** The authors declare that they have no competing interests. **Data and materials availability:** All data needed to evaluate the conclusions in the paper are present in the paper and/or the Supplementary Materials. Additional data related to this paper may be requested from the authors.

Submitted 8 October 2016

Accepted 2 February 2017

Published 8 March 2017

10.1126/sciadv.1602467

**Citation:** S. Liu, W. Yao, Y. Liu, Q. Wei, J. Chen, X. Wu, F. Xia, W. Hu, A Rh(II)-catalyzed multicomponent reaction by trapping an  $\alpha$ -amino enol intermediate in a traditional two-component reaction pathway. *Sci. Adv.* **3**, e1602467 (2017).

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*Sci Adv* 3 (3), e1602467.  
DOI: 10.1126/sciadv.1602467

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