CHEMICAL PHYSICS

Using parahydrogen to hyperpolarize amines, amides, carboxylic acids, alcohols, phosphates, and carbonates

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Hyperpolarization turns weak nuclear magnetic resonance (NMR) and magnetic resonance imaging (MRI) responses into strong signals, so normally impractical measurements are possible. We use parahydrogen to rapidly hyperpolarize appropriate $^1$H, $^{13}$C, $^{15}$N, and $^{31}$P responses of analytes (such as NH$_3$) and important amines (such as phenylethylamine), amides (such as acetamide, urea, and N-methacrylamide), alcohols spanning methanol through octanol and glucose, the sodium salts of carboxylic acids (such as acetic acid and pyruvic acid), sodium phosphate, disodium adenosine 5'-triphosphate, and sodium hydrogen carbonate. The associated signal gains are used to demonstrate that it is possible to collect informative single-shot NMR spectra of these analytes in seconds at the micromole level in a 9.4-T observation field. To achieve these wide-ranging signal gains, we first use the signal amplification by reversible exchange (SABRE) process to hyperpolarize an amine or ammonia and then use their exchangeable NH protons to relay polarization into the analyte without changing its identity. We found that the $^1$H signal gains reach as high as 650-fold per proton, whereas for $^{13}$C, the corresponding signal gains achieved in a $^1$H-$^{13}$C refocused insensitive nuclei enhanced by polarization transfer (INEPT) experiment exceed 570-fold and those in a direct-detected $^{13}$C measurement exceed 400-fold. Thirty-one examples are described to demonstrate the applicability of this technique.

INTRODUCTION

Nuclear magnetic resonance (NMR) is one of the most powerful methods for the study of materials, and magnetic resonance imaging (MRI) plays a vital role in clinical diagnosis. However, the low sensitivity of these techniques limits their applicability. The hyperpolarization method dynamic nuclear polarization (DNP) improves the detectability of analytes such as pyruvate to the level that the MRI-based diagnosis of disease is now possible (1). Parahydrogen ($^2$H$_2$), which is cheap to prepare and exists as a pure nuclear spin state, was shown to enhance the strength of an NMR signal in 1987 (2), although these methods have not yet been used clinically. This may reflect the fact that $^2$H$_2$ was originally used to sensitize chemically modified hydrogenation products (3, 4), and only recently has a method been developed where the original identity of the sensitized analyte is retained (5). This approach, signal amplification by reversible exchange (SABRE), harnesses $^2$H$_2$ in the form of metal-bound hydride ligands and transfers hyperpolarization into a weakly bound substrate (6–8) via the small J-couplings that connect them (9). Ligand exchange then builds up a pool of hyperpolarized substrate according to Scheme 1A (10). SABRE is successful for analytes with multiple bonds to nitrogen such as nicotinamide (11), isoniazid (12), pyrazole (13), and acetonitrile (14), with $^1$H polarizations of 50% (11) and $^{15}$N values of 20% (15) being achieved. Furthermore, although it works for other nuclei (11, 16–20), it fails to sensitize many classes of analytes.

Here, we describe a method where $^2$H$_2$ hyperpolarizes a range of amines, amides, carboxylic acids, alcohols, phosphates, and carbonates without changing their chemical identity. Our method starts with the hyperpolarization of ammonia (the hyperpolarization transfer agent). Subsequently, polarization is relayed into the specified analyte through proton exchange, as outlined in Scheme 1B. Spontaneous low-field transfer then creates the hyperpolarized analyte, which we detect. We called this approach SABRE-RELAY and predict that, when it is fully optimized, it will have a major impact on NMR and MRI in accordance with the fact that we exemplify it for 31 analytes.

RESULTS

We achieve SABRE-RELAY by reacting ammonia with the most versatile of the current SABRE catalysts, [IrCl(COD)(IMes)]$_2$ (21, 22) (1) [where IMes is 1,3-bis(2,4,6-trimethylphenyl)imidazol-2-ylidine and COD is cycloocta-1,5-diene] and H$_2$ to form [Ir(H$_2$)(IMes)(NH$_3$)$_3$]Cl (2-NH$_3$) according to Scheme 1. When this reaction is competed in dichloromethane-$d_2$, 2-NH$_3$ exhibits equatorial and axial NH$_3$ ligand signals at δ 2.19 and 2.88 in the corresponding $^1$H NMR spectrum, alongside a broad NH$_3$ response at δ 0.47, as detailed in Fig. 1A. When this sample is examined after exposure to a 2-bar pressure of $^2$H$_2$ gas at 60 G, the resulting $^1$H NMR signal for free NH$_3$ now shows an ~10-fold signal enhancement per proton, with the bound NH$_3$ ligand signal at δ 2.19 showing a 3-fold enhanced response. These observations confirm that 2-NH$_3$ undergoes SABRE to produce hyperpolarized ammonia. When the same process is repeated in methanol-$d_4$, 2-NH$_3$ exhibits a hydride resonance at δ −23.2 that rapidly separates into several components as H-D exchange proceeds to form an array of isotopologues. However, when $^2$H$_2$–H$_2$ is used, a hyperpolarized NMR signal is readily seen at δ 5.06 for the exchangeable proton of CD$_3$OH, which exhibits a 32-fold intensity gain over its thermally equilibrated signal. Therefore, we added a 5% loading of H$_2$O, relative to iridium, to the CD$_2$Cl$_2$ sample and reexamined it. Under these conditions, the free NH$_3$ signal gain resulting from SABRE proved to increase to 40-fold per proton, whereas the corresponding equatorial ligand signal now showed an 85-fold per proton gain (Fig. 1B). In addition, the free H$_2$O signal was enhanced by 75-fold per proton, a result that compares well with other solvent signal enhancements (23–25).

Exchange spectroscopy measurements were then used to confirm that free NH$_3$ and the equatorially bound NH$_3$ ligand of 2-NH$_3$ are in chemical exchange, with the observation of further exchange peaks between free NH$_3$ and H$_2$O demonstrating the rapid transfer of protons between them. On the basis of this selectivity, we conclude that, when the ammonia is bound, proton exchange between NH$_3$ and H$_2$O is suppressed because the nitrogen lone pair is involved in bonding to the metal center. Consequently, it now becomes hyperpolarized by SABRE. Proton exchange proceeds, though, after NH$_3$ dissociation, and this leads to the observation of hyperpolarization in the chemical exchange–averaged...
response of H2O (or HOCD3) according to Scheme 1B. Now, we show how it is possible to harness this proton exchange process to hyperpolarize the NMR signals of a series of added analytes.

First, we consider whether the SABRE hyperpolarization of NH3 can be relayed into the 1H and 13C responses of a series of alcohols CH3(CH2)nOH (where n = 0 to 7). To do this, we prepared a range of dichloromethane-d2 solutions that contained [Ir(H)2(IMes)(NH3)3]Cl (2-NH3), NH3, and 1 µl of each alcohol (typical concentration, 20 mM). After hyperpolarization transfer from p-H2, strong signals resulted in the associated single-scan 1H NMR spectra, which reached up to 650-fold intensity gains per alcohol CH proton for 1-propanol, averaging at 265 across the series (see the Supplementary Materials). When the same p-H2 transfer process was undertaken and a fully coupled 13C NMR measurement was made instead of a 1H NMR measurement, molecule-diagnostic 13C and 1H,13C refocused insensitive nuclei enhanced by polarization transfer (INEPT)–based responses could also be recorded in one scan at 9.4 T for all the alcohols, as illustrated in Fig. 2B for 1-pentanol, with the associated signal gains reaching 570-fold for the Cα signal of 1-hexanol. The SABRE-RELAY effect results in the detection of hyperpolarized NMR signals for all the spin-1/2 nuclei in these molecules. In addition, as with SABRE, the hyperpolarized NMR terms reflect a mixture of longitudinal single-spin and higher-order states, whose relative amplitudes depend on the magnetic field that the sample experiences during the polarization transfer step (16, 26). Furthermore, by reducing the concentrations of these analytes below the concentration of NH3, it is possible to improve on SABRE-RELAY efficiency. This is beneficial when studying low-concentration analytes because

Fig. 1. Hyperpolarization of NH3 under SABRE. (A) Thermally polarized control 1H NMR spectrum showing peaks for 2-NH3, NH3, and H2 at 298 K in dichloromethane-d2, ×128 vertical expansion relative to (B). (B) Corresponding single-scan 1H NMR spectrum in the presence of p-H2, with the hyperpolarized responses for H2O, NH3(eq), Ir-NH3(eq), and Ir-NH3(axial) of 2-NH3 indicated.
when propanol was studied, the $^1H$ NMR signal gains seen for its OH resonance increased by 100% on moving from a 15 to 1.5 mM concentration (fig. S4), whereas its CH resonances showed a ca. 50% improvement in enhancement level; a $^1H-^{13}C$ refocused INEPT response was still clearly visible in one scan, where the three signals from the OH end were 639, 538, and 603 times larger, respectively, than those in the corresponding $^{13}C$ response. This polarization transfer method is also applicable to complex branched alcohols, and when a sample of $^{13}C$-labeled glucose was analyzed, a single-scan $^{13}C$ response could be seen for all the expected $\alpha$ and $\beta$ form signals, which serves to illustrate the wider significance of this effect (fig. S15E). Furthermore, our studies show that, when SABRE-RELAY is carried out under anhydrous conditions with straight-chain alcohols, superior results are obtained.

Our next goal was to expand on the range of materials that can be sensitized by this method. We started with pyruvic acid but found that its addition to a solution of $2-NH_2$ and $p-H_2$ resulted in ammonium salt precipitation, which acted to limit hyperpolarization efficacy. This can be overcome by the addition of a pH modifier such as Cs$_2$CO$_3$, but working with the corresponding sodium salt proved optimal. When $^{13}C$-labeled sodium pyruvate, acetate, or propionic acid samples were studied in the presence of $p-H_2$, strong $^1H$ and $^{13}C$ signals were seen; the $^{13}C$ signal gain for propionic acid was 109-fold. Furthermore, sodium dihydrogen phosphate, adenosine 5'-triphosphate disodium, and $^{13}C$-labeled sodium hydrogen carbonate provided strong $^{31}P$ and $^{13}C$ responses (Fig. 3, A and B), whereas the amides acetamide, urea, and methacrylamide showed substantial $^1H$, $^{13}C$, and $^{15}N$ signal gains; for urea, a $^{13}C$ signal gain of 408-fold was observed. These studies could be completed with NH$_2$H$_2$O or NH$_2$CH$_2$OH, as detailed in the Supplementary Materials, to promote the necessary proton exchange, and the observations establish that analytes containing the four common functional groups—OH, NH$_2$, CO, and COOH—can be used. In some cases, we see evidence for Schiff-base condensation at long reaction times but could suppress this process by adding water.

To examine the role of the hyperpolarization transfer agent, we replaced NH$_3$ with benzylamine (BnNH$_2$) or phenethylamine (PEA).

Both react with I and H$_2$, forming [Ir(H)$_2$(IMes)(NH$_2$Bn)$_3$]Cl (2-BnNH$_2$) and [Ir(H)$_2$(IMes)(PEA)$_3$]Cl (2-PEA), respectively. For the corresponding 2-BnNH$_2$ sample, signal gains for free BnNH$_2$ of 72-fold (NH), 53-fold (CH), and 170-fold (aromatic), respectively, per proton are observed (Fig. 4), and these measurements can be repeated if the same sample is probed with $p-H_2$ several days after the first observation was made. PEA proved to perform better than BnNH$_2$, with the corresponding NH$_2$ signal gain being 108-fold per proton for a 10-fold loading of 1 with signal gains of 50-fold (NCH$_2$), 45-fold (CH$_2$), 92-fold (ortho), 50-fold (meta), and 20-fold (para) resulting for the other groups. These observations show how polarization transfer through the aliphatic carbon chain into the aromatic protons is possible. BnNH$_2$ and PEA also proved suitable for SABRE-RELAY. In the case of PEA, the efficiency of urea hyperpolarization was found to improve (Fig. 3, C and D) over that achieved with NH$_3$, although the measured response of $^{13}C$-labeled glucose was found to reduce. Furthermore, replacing BnNH$_2$ with its $d_7$-form, C$_6$D$_5$CD$_2$NH$_2$, led to further improvements in observed analyte response level because the initially created SABRE hyperpolarization was now optimally focused into just the NH$_2$ protons.

Given the wide range of amine $p_K_a$ values (27), it may be possible to remove the need for an auxiliary base when dealing with acidic analytes through a process of amine variation. Therefore, we conclude that the facts about the role of the amine will be important for the optimization of SABRE-RELAY and may even allow the introduction of selectivity into the hyperpolarization process. Furthermore, because improvements in analyte detectability with SABRE can be easily achieved by varying the polarization transfer field, reducing relaxation within the analyze, and optimizing the catalyst lifetime while minimizing its reactivity, we expect the signal gains reported here to be similarly improved upon in the future (5).

**DISCUSSION**

In summary, we have shown that SABRE-RELAY can be used to hyperpolarize a wide range of biologically relevant materials. In the initial
step, SABRE is used to enhance the NH proton response of the selected hyperpolarized transfer agent (the free amine) by between 10- and 120-fold per proton. When this is achieved in the presence of propanol, proton exchange results in its OH signal being amplified by between 250- and 500-fold. The nonequilibrium magnetic state of the OH proton is then successfully relayed into its aliphatic 1H resonances such that the corresponding signals are amplified by between 650- and 790-fold per proton. We used this 1H signal gain to record a single-scan 1H-13C spectrum of hyperpolarized BnNH2 achieved via 2-BnNH2 (top right) under SABRE in dichloromethane-d2 solution after transfer at 60 °C (see fig. S32A for the corresponding thermally equilibrated NMR spectrum). The enhanced signals are color-coded; NH2 (blue), CH2 (light green), and Ph (orange) for the bound equatorial BnNH2 ligand of 2-BnNH2, NH2 and CH2 of the free material, and H2O. (B) Corresponding 15N NMR spectrum recorded using 15N-labeled BnNH2 after transfer in a μ-metal shield showing the free (left) and equatorial ligand (right) responses.

**SUPPLEMENTARY MATERIALS**

Supplementary material for this article is available at http://advances.sciencemag.org/cgi/content/full/4/1/eaa06250/DC1

section S1. SABRE-RELAY polarization transfer method with NH3

section S2. SABRE-RELAY polarization transfer method with BnNH2 or PEA

section S3. Polarization enhancement quantification procedures

section S4. NMR spectrometer details

section S5. Pulse sequence details

section S6. SABRE-RELAY spectra

fig. S1. INEPT pulse sequence.

fig. S2. DEPT pulse sequence.

fig. S3. SABRE-RELAY NMR spectra methanol.

fig. S4. SABRE-RELAY NMR spectra ethanol.

fig. S5. SABRE-RELAY NMR spectra propanol.

fig. S6. SABRE-RELAY NMR spectra low concentration.

fig. S7. SABRE-RELAY NMR spectra pentanol.

fig. S8. SABRE-RELAY NMR spectra hexanol.

fig. S9. SABRE-RELAY NMR spectra heptanol.

fig. S10. SABRE-RELAY NMR spectra octanol.

fig. S11. SABRE-RELAY NMR spectra tert-butanol.

fig. S12. SABRE-RELAY NMR spectra iso-propanol.

fig. S13. SABRE-RELAY NMR spectra tert-butanol.

fig. S14. SABRE-RELAY NMR spectra Δ-glucose.

fig. S15. SABRE-RELAY NMR spectra α-glucose-13C.

fig. S16. SABRE-RELAY NMR spectra glycero.

fig. S17. SABRE-RELAY NMR spectra sodium acetate-13C.

fig. S18. SABRE-RELAY NMR spectra sodium pyruvate-13C.

fig. S19. SABRE-RELAY NMR spectra sodium acetate-1,2 13C2.

fig. S20. SABRE-RELAY NMR spectra propionic acid-13C.

fig. S21. SABRE-RELAY NMR spectra sodium hydrogen carbonate-13C.

fig. S22. SABRE-RELAY NMR spectra urea-13C.

fig. S23. SABRE-RELAY NMR spectra urea-15N2.

fig. S24. SABRE-RELAY NMR spectra urea-13C-15N2.

fig. S25. SABRE-RELAY NMR spectra acetylamide.

fig. S26. SABRE-RELAY NMR spectra methachloramide.

fig. S27. SABRE-RELAY NMR spectra cyclohexyl methachloramide.

fig. S28. SABRE-RELAY NMR spectra mono sodium dihydrogen orthophosphate.
fig. S29. SABRE-RELAY NMR spectra adenosine 5'-triphosphate disodium salt.

fig. S30. SABRE-RELAY NMR spectra ammonia in dichloromethane.

fig. S31. SABRE-RELAY NMR spectra ammonia in methanol.

fig. S32. SABRE-RELAY NMR spectra benzylamine.

table S2. Alcohol $^{13}$C SABRE-RELAY signal enhancement values.

REFERENCES AND NOTES


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and design of experiments. Competing interests: W.I., P.J.R., and S.B.D. are inventors on a patent application filed by the University of York based on this work (patent no. GB 1711967.8, filed 25 July 2017). The authors declare that they have no other competing interests. Data and materials availability: All data needed to evaluate the conclusions in the paper are present in the paper and/or the Supplementary Materials. All relevant plasmids and experimental data may be requested from the authors. Raw data can be found via DOI: 10.15124/d93f83bf-9fc7-4cfb-8267-504a0304a516.

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